SNC 2P - 2020-05-22 (Friday)

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Date: Fri, 22 May 2020 15:25:16 +0000 (*2020-05-22 11:25:16 AM*)

Hello All:

Hope you are well!

Please Check In: https://forms.office.com/Pages/ResponsePage.aspx?id=GAmpRLReCU2WCd35yhGvQuASPjs6aYVFk-EAh60FvohUN0VTRTZUNlBDRThNRUtENks3QkVJSUVVOCQlQCN0PWcu

The goal today is to learn about atoms:

Please look at this note http://www.schlenkerchem.org/2P/Chemistry/Notes/note%203%20-%20atoms%20and%20pt.pdf

AND WHILE YOU ARE ANSWERING THE QUESTIONS USE THIS PERIODIC TABLE: https://www.ptable.com/

As you know, the periodic table is an organized arrangement of the known elements. Please answer these questions as you go through the note:

From the table at the top of the note and the first paragraph:

1. What is the name and symbol for the positive subatomic particle and where is it found?
2. What is the name and symbol for the neutral subatomic particle and where is it found?
3. What is the name and symbol for the negative subatomic particle and where is it found?
4. Which of the above three particles has considerably less mass than the other two?
5. Which of the above subatomic particles determines the type of atom (element) that you have?

From the periodic table:

6. Using the periodic table, state how many protons are required to build a nucleus for:

> hydrogen (h)

> lithium (Li)
>
> carbon (C)
>
> oxygen (O)

sodium (Na)

iron (Fe)

uranium (U)

From the second paragraph in the note and the diagram on the second page - scroll down

7. What did Neils Bohr find out?

8. From the diagram:

How many electrons are in the inner most circle?

How many electrons are in the next circle?

How many electrons are in the outermost circle and how many more electrons could you add to this circle to make it full?

9. Add up the total number of electrons in the diagram and compare it to the number of protons. What do you notice and what does this mean?

Have a good Weekend!

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