

## Further comments

**From:** Fred Schlenker <fred\_schlenker@bwdsb.on.ca>

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The cell potential at the side of the half reactions table uses a standard hydrogen half cell as the zero point (concept of zero point was made in the atomic structure unit and the thermodynamics unit). Just like Hess' law, flipping an equation will flip the sign on the potential energy. Unlike Hess' Law, when you multiply an equation you do not multiply the potential. The difference is, that in thermodynamics the enthalpy values represent a quantity of heat, where as the potential represents a tendency to push/pull electrons from one half cell to another. If you think of a house sized battery that is capable of generating 9 V of electricity, the size of the battery does not change the potential energy behind the electrons. It can still only run a flash light.

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