Redox Reactions: Part #3 - Balancing Redox Equations Using Standard Half Cell Reactions - SCH 4U Independent Study Unit

Use the table found in your text book on page 805. For each of the following half cell pairs, write a balanced redox reactions such that the reaction is written in the spontaneous direction. Add cell voltages to determine the overall potential of REDOX reaction.

a) the three electron reaction involving gold with eight electron reaction involving perchlorate

b) the two electron reaction of copper with a one electron reaction involving iron

c) the two electron reaction of zinc with the one electron reaction involving nitrate $% \left(1\right) =\left(1\right) \left(1\right) +\left(1\right) +\left(1\right) \left(1\right) +\left(1\right$