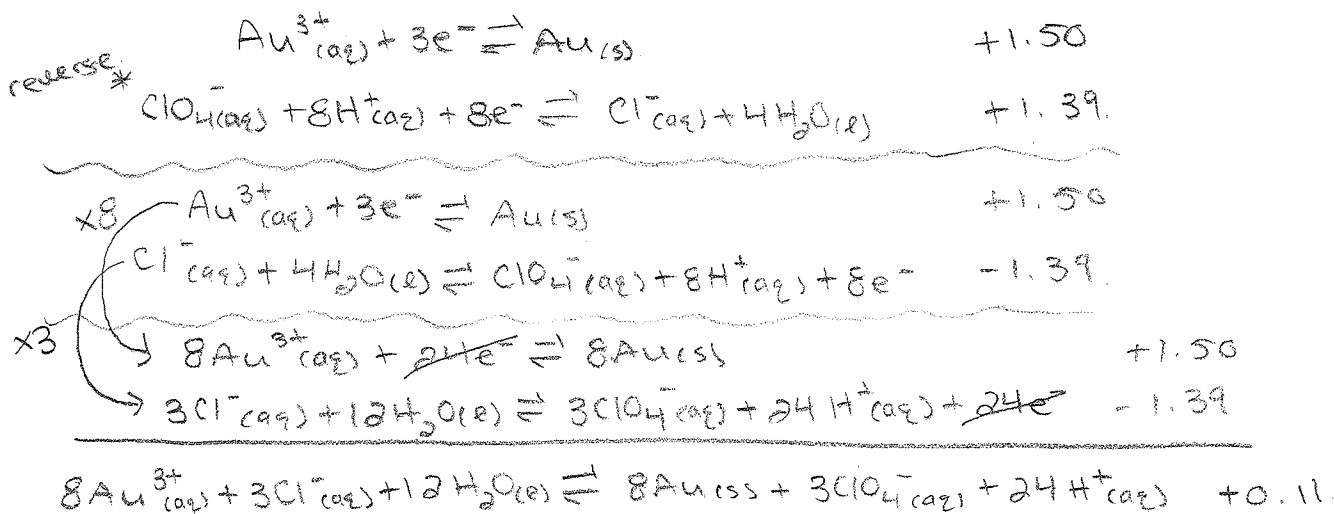


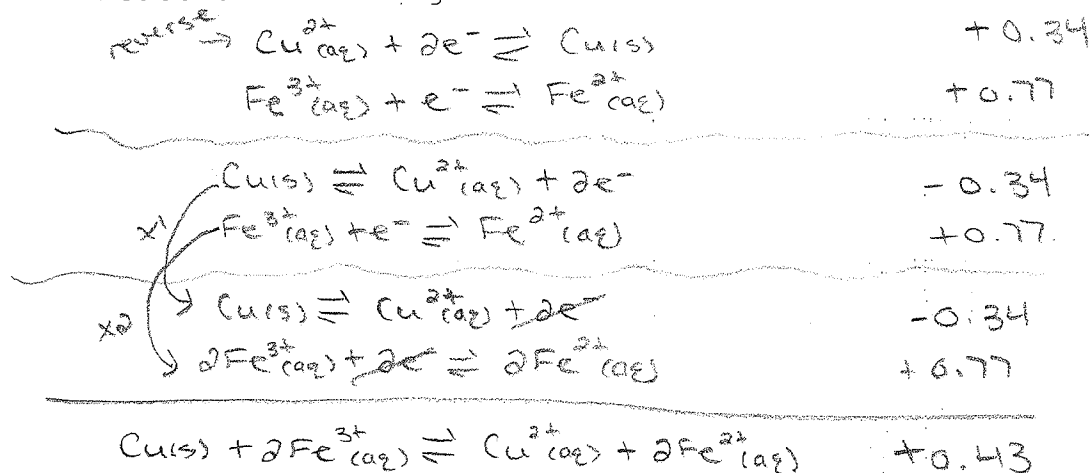
Redox Reactions: Part #3 - Balancing Redox Equations Using Standard Half Cell Reactions - SCH 4U Independent Study Unit

Use the table found in your text book on page 805. For each of the following half cell pairs, write a balanced redox reactions such that the reaction is written in the spontaneous direction.

- a) the three electron reaction involving gold with eight electron reaction involving perchlorate



- b) the two electron reaction of copper with a one electron reaction involving iron



- c) the two electron reaction of zinc with the one electron reaction involving nitrate

