Redox Reactions: Part #1 - Oxidation States SCH 4U Independent Study Unit

Practice Quiz:

For each compound, write the oxidation state for every element directly above the element. I may be useful to consider the overall ionic charge for polyatomic anions. It may also be necessary to consider structures of organic compounds!

$\overset{1+}{K}\overset{7+}{Mn}\overset{2-}{O}_4$

H^{1+2+1-}

$$\operatorname{Fe}_{2} \left(\begin{array}{c} 6+2-\\ \mathbf{S} \mathbf{O}_{4} \end{array} \right)_{3}$$

${{\rm Cu}\,{\rm CN}}^{1+2+3-}$

$\overset{3-}{C}\overset{1+}{H}_3\overset{3+}{C}\overset{2-}{O}\overset{2-}{O}\overset{1+}{H}$