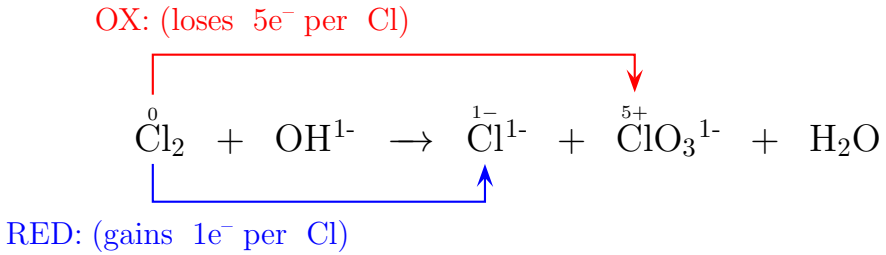
EXAMPLE OF IONIC EQUATION

24. $\overset{0}{Cl}_2 + OH^{1-} \rightarrow \overset{1-}{Cl}^{1-} + \overset{5+}{Cl}O_3{}^{1-} + H_2O$

1. Assign oxidation states.

EXAMPLE OF IONIC EQUATION



1. Assign oxidation states.
2. Identify the losers and gainers.

EXAMPLE OF IONIC EQUATION

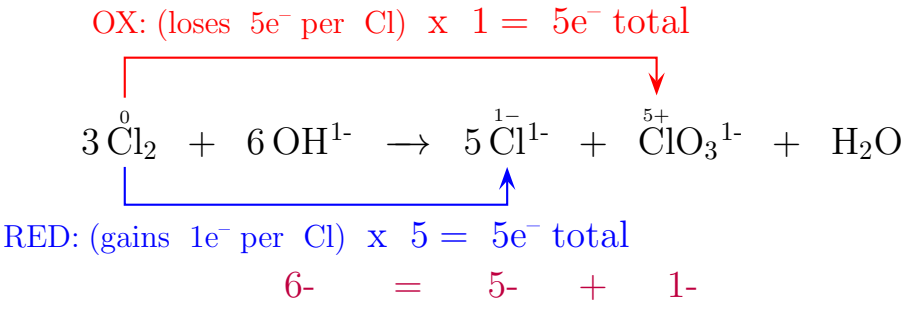
OX: (loses $5e^-$ per Cl) x 1 = $5e^-$ total

24. $3\,\overset{0}{Cl_2} + OH^{1-} \rightarrow 5\,\overset{1-}{Cl}^{1-} + \overset{5+}{ClO_3}^{1-} + H_2O$

RED: (gains $1e^-$ per Cl) x 5 = $5e^-$ total

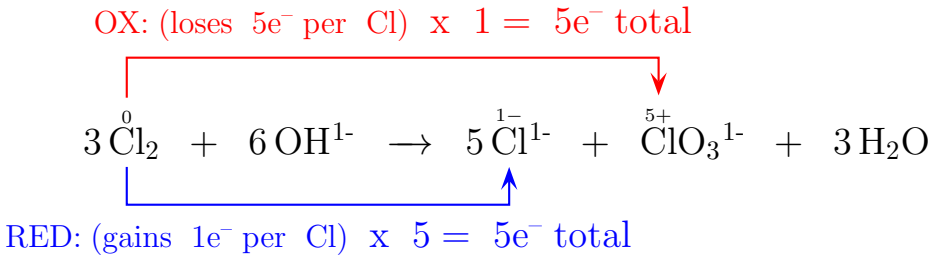
1. Assign oxidation states.
2. Identify the losers and gainers.
3. Balance the number of electrons lost and gained (pay attention to any additional stoichiometric considerations).

EXAMPLE OF IONIC EQUATION



1. Assign oxidation states.
2. Identify the losers and gainers.
3. Balance the number of electrons lost and gained (pay attention to any additional stoichiometric considerations).
4. Balance the remainder of the equation using:
   - charge balance if ions are present

EXAMPLE OF IONIC EQUATION



1. Assign oxidation states.
2. Identify the losers and gainers.
3. Balance the number of electrons lost and gained (pay attention to any additional stoichiometric considerations).
4. Balance the remainder of the equation using:
   - charge balance if ions are present
   - other atoms
   - hydrogen
   - oxygen