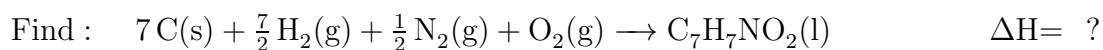


HESS' LAW 6

$$\text{conversion for CO}_2(\text{g}): -393.5 \text{ kJ} \times \frac{1 \text{ kcal}}{4.184 \text{ kJ}} = -94.0 \text{ kcal}$$

$$\text{conversion for H}_2\text{O}(\text{g}): -285.8 \text{ kJ} \times \frac{1 \text{ kcal}}{4.184 \text{ kJ}} = -68.3 \text{ kcal}$$

$$\text{conversion for NO}_2(\text{g}): +33.2 \text{ kJ} \times \frac{1 \text{ kcal}}{4.184 \text{ kJ}} = +7.93 \text{ kcal}$$



Given :

