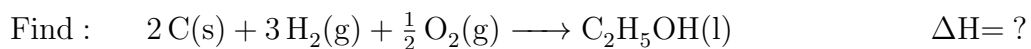


HESS' LAW 4

conversion for $\text{CO}_2(\text{g})$: $-393.5 \text{ kJ} \times \frac{1 \text{ kcal}}{4.184 \text{ kJ}} = -94.0 \text{ kcal}$

conversion for $\text{H}_2\text{O}(\text{g})$: $-285.8 \text{ kJ} \times \frac{1 \text{ kcal}}{4.184 \text{ kJ}} = -68.3 \text{ kcal}$



Given :

