

Equilibrium Expression

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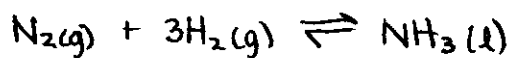
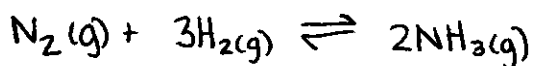
- sometimes called Equilibrium Law



$$\left[K_{eq} = \frac{[C]^c [D]^d}{[A]^a [B]^b} \right] \quad \leftarrow \text{equilibrium "expression"}$$

↳ only substances with 'variable' concentration are included in the equilibrium expression. (g)(aq)

ex



$$K_{eq} = \frac{[NH_3]^2}{[N_2][H_2]^3}$$

$$K_{eq} = \frac{1}{[N_2][H_2]^3}$$

K_{eq} - is a number

- is called the equilibrium constant

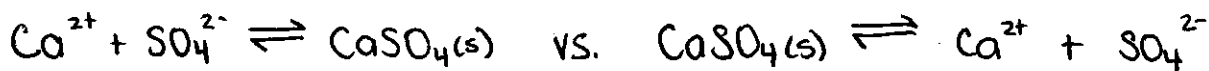
- can be looked up in tables

- temperature dependant (value changes for different temperatures)

- is not affected by anything else

- is dependant on the way the equilibrium is written

ex



$$K_{eq} = \frac{1}{[Ca^{2+}][SO_4^{2-}]}$$

$$K_{eq} = [Ca^{2+}][SO_4^{2-}]$$

↖ preferred choice