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Initial Amount			
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400 mL + 500 mL = 900 mL

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Final Amount	1.137×10^{-5} mol		
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$$2.244\,mg\ BaCO_3 \times \frac{1\,g}{1000\,mg} \times \frac{1\,mol\ BaCO_3}{197.34\,g\ BaCO_3} = 1.137 \times 10^{-5}\ mol\ BaCO_3$$

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Final Amount	1.137×10^{-5} mol	$4 \times 10^{-5} - 1.137 \times 10^{-5} = 2.863 \times 10^{-5}$ mol	$1.5 \times 10^{-4} - 1.137 \times 10^{-5} = 1.386 \times 10^{-4}$ mol
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Final [ ]	N.A.	3.181×10^{-5} M	$1.5\,40 \times 10^{-4}$ M

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Final Amount	1.137×10^{-5} mol	$4 \times 10^{-5} - 1.137 \times 10^{-5} = 2.863 \times 10^{-5}$ mol	$1.5 \times 10^{-4} - 1.137 \times 10^{-5} = 1.386 \times 10^{-4}$ mol
Final [ ]	N.A.	$3.181 \times 10^{-5}\,M$	$1.5\,40 \times 10^{-4}\,M$

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$$Ksp = [Ba^{2+}][CO_3^{2-}]$$

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Final Amount	1.137×10^{-5} mol	$4 \times 10^{-5} - 1.137 \times 10^{-5} = 2.863 \times 10^{-5}$ mol	$1.5 \times 10^{-4} - 1.137 \times 10^{-5} = 1.386 \times 10^{-4}$ mol
Final [ ]	N.A.	$3.181 \times 10^{-5}\,M$	$1.5540 \times 10^{-4}\,M$

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$$K_{sp} = [Ba^{2+}][CO_3^{2-}]$$

$$K_{sp} = (3.181 \times 10^{-5})(1.540 \times 10^{-4})$$

$$K_{sp} = 4.9 \times 10^{-9}$$