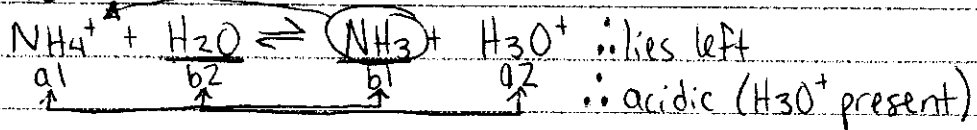


\*SPECIAL CASE

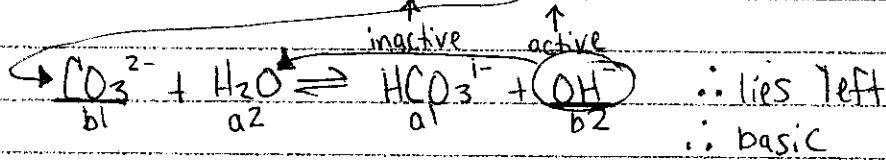
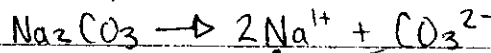


- active anions is any anion that is the conjugate base of an acid that is weaker than hydronium

$\therefore \text{IO}_3^-, \text{HSO}_3^-, \text{SO}_4^{2-}, \text{H}_2\text{PO}_4^-$  etc. all are increasingly active

$\therefore \text{ClO}_4^-, \text{I}^-, \text{Br}^-, \text{Cl}^-, \text{NO}_3^-, \text{HSO}_4^-, \text{H}_2\text{O}$  are inactive  
 $\text{SO}_4^{2-}$  exception

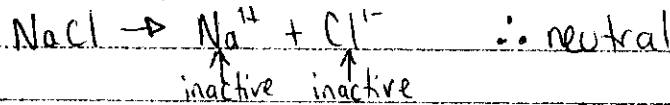
• eg: active anion, inactive cation



- active anions lead to basic

- inactive anions lead to neutral (as in, no reaction)

• eg: inactive anion + inactive cation



• eg: active anion + active cation

- truly evil & not part of this course

- problem = low solubility

- comparison of cation & anion equilibria required = unpleasant!