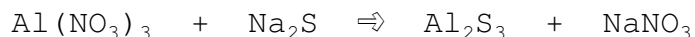


**Stoichiometry Using Solutions**  
**SCH 4C**

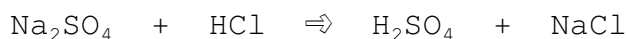
1. Determine the mass of aluminum sulphide that will form when 450 mL of 0.2 M sodium sulphide is reacted.



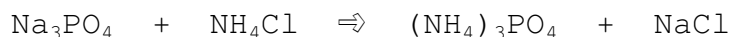
2. Determine the concentration of silver nitrate that would form from the reaction of 0.400 g of silver. The total volume of the solution is 250 mL



3. What concentration of sulphuric acid results from the reaction of 50.0 mL of 0.25 M sodium sulphate. Total volume of resulting solution will be 75.0 mL



4. What was the concentration of 125 mL of sodium phosphate if the resulting concentration of ammonium phosphate is 3.5 M for 250 mL of solution.



5. 100 mL of 0.1 M lead(II) nitrate is mixed with 100 mL of 0.1 M sodium chloride. Determine the maximum possible mass of lead(II) chloride that can form.

