

### SCH 4C In Class Conversion Assignment

Write complete solutions for the following questions from your text book. All questions are from the practice problems:

pg 94 #1, 2

pg 95 #3, 4, 5

pg 98 #6, 7, 8

pg 100 #12, 13

pg 102 #14, 15

pg 103 #16, 17(formula unit is the same as a "molecule")

pg 105 #18, 19, 20, 21

**Please note that the answers for each question is given in the margin. Therefore there are no marks for a correct answer. All marks are for the format that you have used to arrive at the answer. Marks will be deducted for poor format or difficult to read material. Do not use the formats found in the text book. The following is a clear example of the format that we have learned. You must follow this format for full marks. For now it is a requirement to show molar mass calculations for all formula (even formula like N<sub>2</sub> or F<sub>2</sub>).**

eg) Determine the number of atoms present in 34.6 g of sodium carbonate (Na<sub>2</sub>CO<sub>3</sub>):

Na: 2 x 22.99 = 45.98 g

C: 1 x 12.01 = 12.01 g

O: 3 x 16.00 = 48.00 g

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105.99 g

$$34.6 \text{ g Na}_2\text{CO}_3 \times \frac{1 \text{ mol Na}_2\text{CO}_3}{105.99 \text{ g Na}_2\text{CO}_3} \times \frac{6.022 \times 10^{23} \text{ molec Na}_2\text{CO}_3}{1 \text{ mol Na}_2\text{CO}_3} \times \frac{6 \text{ atoms}}{1 \text{ molec Na}_2\text{CO}_3} = 1.18 \times 10^{24} \text{ atoms}$$