Name:

SCH 4C Equation Quiz

- 1. Balance each equation by adding stoichiometric coefficients. Be sure that your answers are in lowest terms. Then identify each reaction as one of:
 - synthesis

 - decomposition
 single replacement
 double replacement
 hydrocarbon combustion
- a) Hg + O_2 \rightarrow Hg₂O type = _____
- Na_3PO_4 + $Mg(NO_3)_2$ \rightarrow $Mg_3(PO_4)_2$ + $NaNO_3$ b) type = _____
- c) $K_2CO_3 \rightarrow K_2O + CO_2$ type = ____
- d) $Sn(SO_4)_2$ + Al \rightarrow $Al_2(SO_4)_3$ + Sntype = _____
- e) C_6H_{14} + O_2 \rightarrow CO_2 + H_2O type = _____
- Al + O_2 \rightarrow Al_2O_3 f) type = _____

- 2. Complete each synthesis reaction and balance:
- a) Sn + N_2 \rightarrow
- b) BaO + H_2O \rightarrow
- 3. Complete each decomposition reaction and balance:
- a) Fe_2O_3 \rightarrow
- b) $CaCO_3$ \rightarrow
- 4. Complete each single replacement reaction and balance:
- a) F_2 + Al_2O_3 \rightarrow
- b) Ca + $Au_2(CO_3)_3$ \rightarrow
- 5. Complete each double replacement reaction and balance:
- a) Na_2SO_4 + $AlCl_3$
- b) $(NH_4)_3PO_4 + SnCl_4 \rightarrow$