

/50 = %

Name: \_\_\_\_\_

**Nomenclature Quiz #3 - SCH 4C**

$C^{4-}$	carbide	$CO_3^{2-}$	carbonate
$N^{3-}$	nitride	$NO_3^{1-}$	nitrate
$O^{2-}$	oxide	$PO_4^{3-}$	phosphate
$F^{1-}$	fluoride	$SO_4^{2-}$	sulphate
$P^{3-}$	phosphide	$ClO_3^{1-}$	chlorate
$S^{2-}$	sulphide	$OH^{1-}$	hydroxide
$Cl^{1-}$	chloride	$CN^{1-}$	cyanide
$As^{3-}$	arsenide		
$Se^{2-}$	selenide	$NH_4^{1+}$	ammonium
$Br^{1-}$	bromide		
$Sb^{3-}$	antimonide		
$Te^{2-}$	telluride		
$I^{1-}$	iodide		

1. Simple monovalent cation (only one oxidation state),  
elemental anion (ends in ide)

- a) NaCl      sodium chloride
- b) Ba<sub>3</sub>N<sub>2</sub>      \_\_\_\_\_
- c) K<sub>2</sub>S      \_\_\_\_\_
- d) Na<sub>3</sub>P      \_\_\_\_\_
- e) CaCl<sub>2</sub>      \_\_\_\_\_
- f) potassium nitride      \_\_\_\_\_
- g) sodium sulphide      \_\_\_\_\_
- h) potassium oxide      \_\_\_\_\_
- i) zinc nitride      \_\_\_\_\_
- j) cadmium chloride      \_\_\_\_\_

2. Polyvalent Cation (more than one possible oxidation state), elemental anion.

1	2	3	4	5	6	7	8	9	10
I	II	III	IV	V	VI	VII	VIII	IX	X

a) arsenic(III) oxide \_\_\_\_\_

b) arsenic(V) oxide \_\_\_\_\_

c) lead(II) sulphide \_\_\_\_\_

d) lead(IV) sulphide \_\_\_\_\_

e) bismuth(V) nitride \_\_\_\_\_

f)  $\text{CuCl}_2$  \_\_\_\_\_

/10

g)  $\text{Sb}_2\text{O}_5$  \_\_\_\_\_

h)  $\text{NiO}$  \_\_\_\_\_

i)  $\text{Fe}_2\text{O}_3$  \_\_\_\_\_

j)  $\text{Mn}_3\text{N}_4$  \_\_\_\_\_

3. Simple monovalent cation with polyatomic anions.

a) sodium hydroxide \_\_\_\_\_

b) calcium cyanide \_\_\_\_\_

c) zinc sulphate \_\_\_\_\_

d) scandium nitrate \_\_\_\_\_

e) aluminum sulphate \_\_\_\_\_

f)  $\text{Na}_2\text{CO}_3$  \_\_\_\_\_

/10

g)  $(\text{NH}_4)_2\text{SO}_4$  \_\_\_\_\_

h)  $\text{Na}_2\text{CO}_3$  \_\_\_\_\_

i)  $\text{AlPO}_4$  \_\_\_\_\_

j)  $\text{Ca}(\text{OH})_2$  \_\_\_\_\_

4. Polyvalent cation with polyatomic ion.

a) carbon(IV) chlorate \_\_\_\_\_

b) iron(II) phosphate \_\_\_\_\_

c) gold(III) sulphate \_\_\_\_\_

d) copper(II) sulphate \_\_\_\_\_

e) copper(I) carbonate \_\_\_\_\_

f)  $\text{Pb}(\text{NO}_3)_4$  \_\_\_\_\_

g)  $\text{Co}_2(\text{SO}_4)_3$  \_\_\_\_\_

h)  $\text{Pb}_3(\text{PO}_4)_4$  \_\_\_\_\_

i)  $\text{Pb}(\text{CO}_3)_2$  \_\_\_\_\_

j)  $\text{Cu}(\text{OH})_2$  \_\_\_\_\_

/10

5. Mixed Problems!!!!

a)  $\text{SO}_3$  \_\_\_\_\_

b)  $\text{K}_3\text{N}$  \_\_\_\_\_

c)  $(\text{NH}_4)_3\text{P}$  \_\_\_\_\_

d)  $\text{ZnCO}_3$  \_\_\_\_\_

e)  $\text{Mn}_2\text{O}_7$  \_\_\_\_\_

f)  $\text{HgSO}_4$  \_\_\_\_\_

g)  $\text{KBr}$  \_\_\_\_\_

h)  $\text{Al}(\text{NO}_3)_3$  \_\_\_\_\_

i)  $\text{Ir}(\text{CO}_3)_3$  \_\_\_\_\_

j)  $\text{ZrS}_2$  \_\_\_\_\_

/10