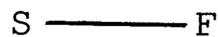
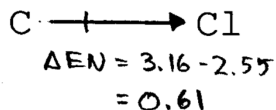


SCH 4C Bond Polarization Quiz

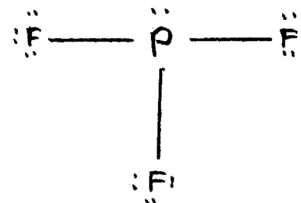
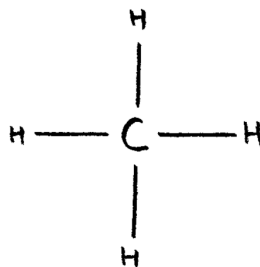
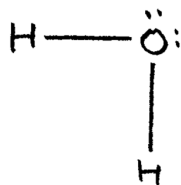
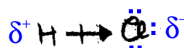
1. For each of the following pairs of atoms, use the electronegativity values to complete a  $\Delta EN$  calculation. Draw in the appropriate bond polarization



/10



2. For each of the following stick structures, add the correct bond polarizations, determine the net molecular polarization and add  $\delta^+$  and  $\delta^-$  as appropriate. If the bond polarizations cancel out, simply write "non-polar"



/6

3. Draw a lewis dot diagram for the covalent bonding you would expect between silicon and chlorine. Then draw the corresponding stick structure and complete as in question #2.

/6