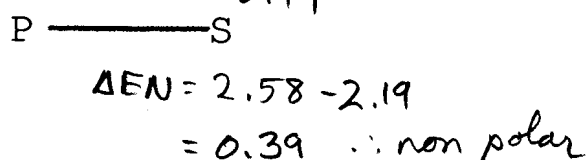
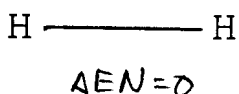
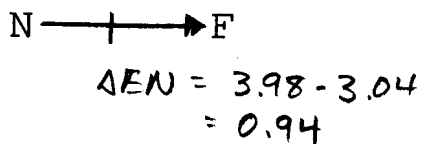
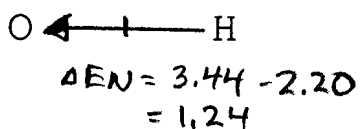
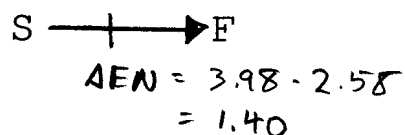
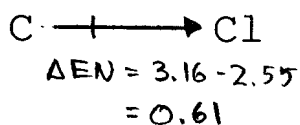
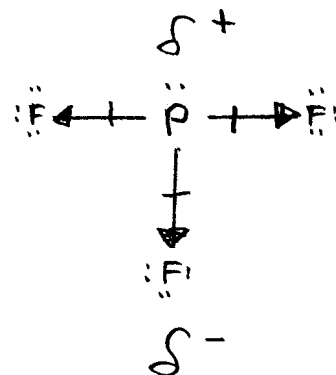
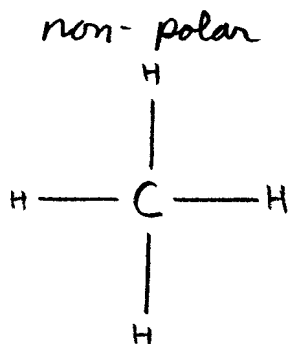
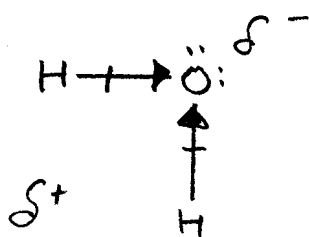
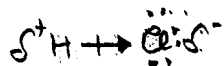


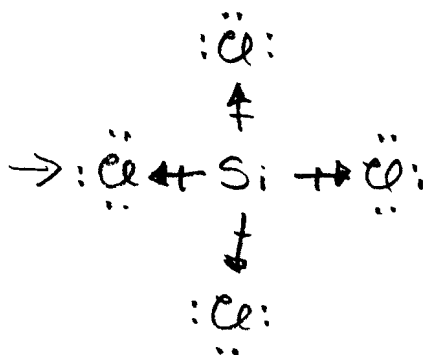
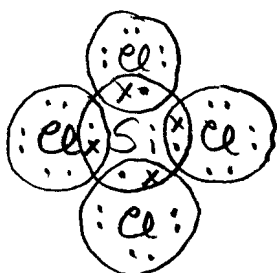
2. For each of the following pairs of atoms, use the electronegativity values to complete a ΔEN calculation. Draw in the appropriate bond polarization



3. For each of the following stick structures, add the correct bond polarizations, determine the net molecular polarization and add δ^+ and δ^- as appropriate. If the bond polarizations cancel out, simply write "non-polar"



4. Draw a Lewis dot diagram for the covalent bonding you would expect between silicon and chlorine. Then draw the corresponding stick structure and complete as in question #2.



non-polar

10

6

6

22