

Name: _____

**Subatomic Particles Atomic Symbols
and Radioactive Decay Quiz - SCH 4C**

Complete this table to show information about all three subatomic particles:

name	symbol	charge	mass	location
proton	p ⁺	1+	1 u*	nucleus
neutron	n	neutral	1 u	nucleus
electron	e ⁻	1-	0.00055 u	orbits nucleus

PLEASE ANSWER IN POINT FORM - MARKING WILL BE EASY - DO YOUR BEST

Why is chemistry more focussed on the behaviour of electrons than the behaviour of protons?

- ELECTRONS CAN CHANGE LOCATION (ADD OR LEAVE AND ATOM)
- PROTONS ARE STUCK IN THE NUCLEUS AND MERELY PROVIDE POSTIVE CHARGE

Why are neutrons considered relatively unimportant in chemistry?

What contribution do neutrons they make to atoms?

- NO CHARGE, THEREFORE DO NOT INTERACT
- ADD MASS TO THE ATOM

Which of the above particles is responsible for determining what type of atom/element you have?

- PROTONS (POSITIVE CHARGE OF THE NUCLEUS DECIDES THE ATOM)

Complete each decay process:

