

MORE CONVERSIONS

- Calculate the amount (moles) in:
 - 25.0 g of oxygen gas
 - 43.8 g of sodium nitrate
 - 0.005 g of potassium bicarbonate
 - 950 g of lithium sulphate
 - 2.2 kg of ammonium phosphate
 - 0.545 g of auric carbonate
 - 3.32 g of potassium perchlorate
 - 0.05 mg of iron(II) nitride
- Calculate the mass in:
 - 1.25 mol of hydrogen oxide
 - 4.3 mol of hydrogen peroxide
 - 0.022 mol of arsenic pentoxide
 - 2.8 mol of carbon dioxide
 - 99.0 mol of nitrogen gas
 - 33.8 mol of neon gas
 - 0.1 mmol of potassium bisulphite
 - 2.2 kmol of lithium hydride
- Calculate the number of molecules in:
 - 2.50 mol of uranium hexafluoride
 - 2.50 mol of ammonium nitrate
 - 2.50 mol of cupric bromide
 - 2.50 mol of stannous chlorate
 - 2.50 mol of tungsten(IV) oxide
 - 2.50 mol of hydrogen chloride
- Calculate the number of atoms in:
 - 2.50 mol of uranium hexafluoride
 - 2.50 mol of ammonium nitrate
 - 2.50 mol of cupric bromide
 - 2.50 mol of stannous chlorate
 - 2.50 mol of tungsten(IV) oxide
 - 2.50 mol of hydrogen chloride
- Calculate the number of molecules in:
 - 3.56 g of lithium oxide
 - 7.89 g of magnesium hypochlorite
 - 89.0 g of sulphur(IV) oxide
 - 0.050 kg of cupric sulphate pentahydrate
 - 21.0 g cobaltous chloride hexahydrate
- Calculate the mass in grams of:
 - 3.50×10^{22} molec of potassium nitrate
 - 8.45×10^{21} molec of silver periodate
 - 1.23×10^{24} molec of sodium thiosulphate pentahydrate ($\text{Na}_2\text{S}_2\text{O}_3 \cdot 5\text{H}_2\text{O}$)
 - 5.45×10^{25} molec of sodium metabisulphite ($\text{Na}_2\text{S}_2\text{O}_5$)
 - 1.11×10^{22} molec of sodium persulphate ($\text{Na}_2\text{S}_2\text{O}_8$)