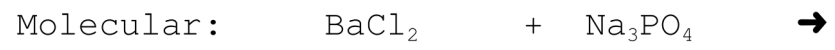


2. barium chloride with sodium phosphate

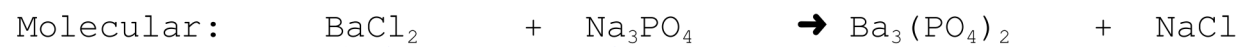


Ionic:

Net Ionic:

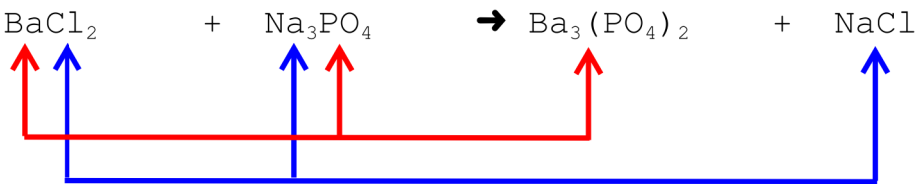
correct reactant formulas

2. barium chloride with sodium phosphate



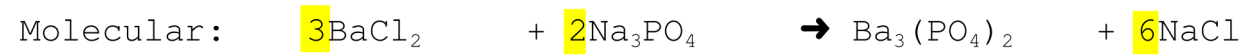
Ionic:

Net Ionic:



double displacement products

2. barium chloride with sodium phosphate



Ionic:

Net Ionic:

balancing

2. barium chloride with sodium phosphate



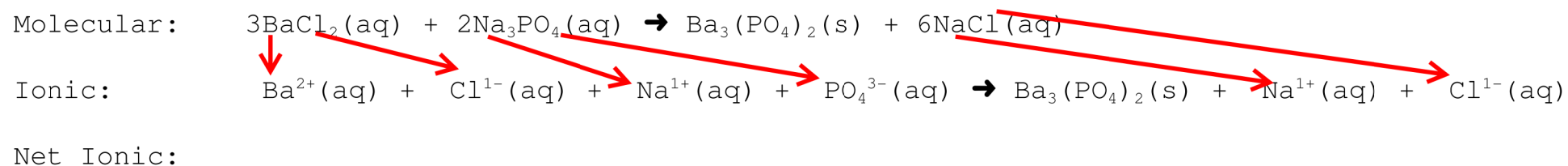
Ionic:

Net Ionic:

adding states

- see solubility rules
- soluble = (aq)
- insoluble = (s)

2. barium chloride with sodium phosphate



show ions

- soluble will break up into a cation and anion (nomenclature)
- insoluble remains together in solid state

DO THIS BEFORE CONSIDERING BALANCING

2. barium chloride with sodium phosphate



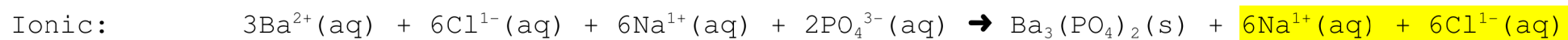
Net Ionic: **balance**

2. barium chloride with sodium phosphate



Net Ionic: **balance**

2. barium chloride with sodium phosphate



Net Ionic:

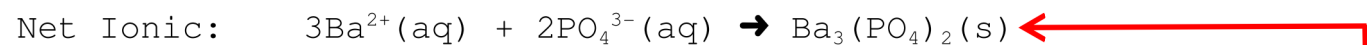
balance

2. barium chloride with sodium phosphate



cancel out ions found on both side
write the leftovers

2. barium chloride with sodium phosphate



therefore barium phosphate is the precipitate

the precipitate (ppte for short) is the solid