/50 = %

Name:_____

SCH 3U – Nomenclature Quiz #1

1. What is rule #1 when doing inorganic nomenclature?

/1

2. What must be true about the total cationic and total anionic charges in order to have a properly constructed chemical formula

/1

3. Classify each of the following elements as elements that form monovalent or polyvalent cations

$_{20}Ca$, $_{26}Fe$, $_{50}Sn$, $_{6}C$, $_{3}Li$

Monovalent	Polyvalent

/2

4. What does the number in roman numerals that follows the cation name tell you exactly? When is it used? When is it not used?

tells you:
used when:
not used when:

/3

5. What are all the possible charges that a $_{17}Cl$ atom can have when it forms ions? Write your answer as ions.

/3

/10

6. Provide either names or formula as appropriate. For polyvalent cation compounds write the I.U.P.A.C. name only:

silver oxide		$SrCl_2$	
carbon(IV) sulphide		SO_2	
magnesium fluoride		Sb_2O_3	
aluminum chloride		Rb_3P	
gallium(III) oxide		Na_4C	
gold(III) nitride		Ca_2C	
copper(II) sulphide		W_2S_5	
vanadium(V) oxide		Hg_2O	
zirconium phosphide		CsF	
mercury(II) bromide		PoO	
thallium(I) oxide		SO_3	
barium arsenide		Tc_3N_4	
platinum(IV) nitride		LaP	
carbon(II) oxide		Mo_3N_5	
carbon(IV) oxide		$PdCl_4$	
hydrogen oxide		CrO_3	
potassium sulphide		$HfSe_2$	
iridium(IV) iodide		ThS_2	
lead(IV) sulphide		XeF_4	
bismuth(V) nitride		$RaSe$	

/20

/20