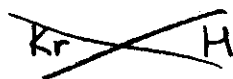
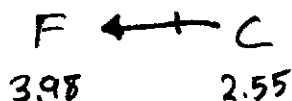
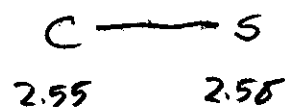
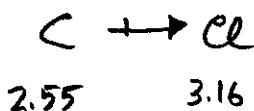
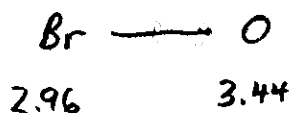
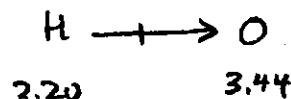
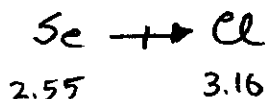
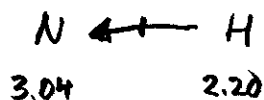
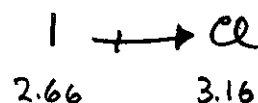
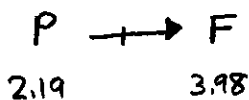
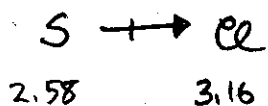
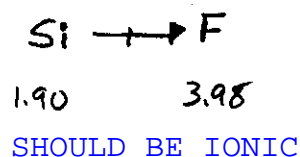
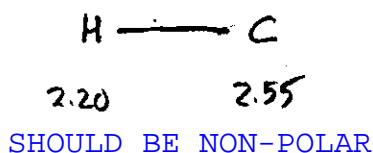
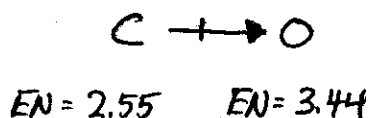


Covalent Molecules

For each of the following pairs of atoms, look up and write down the electronegativity value and use this information to determine the direction of bond polarization. (The first one is done as an example)

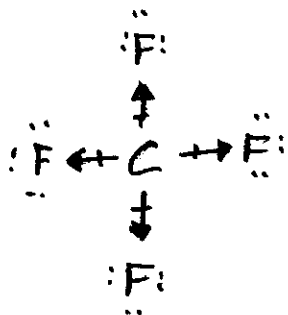


Kr does not bond!

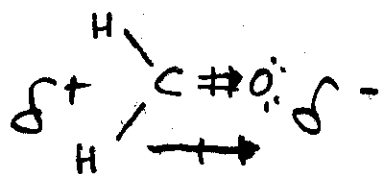


even sharing
no polarization

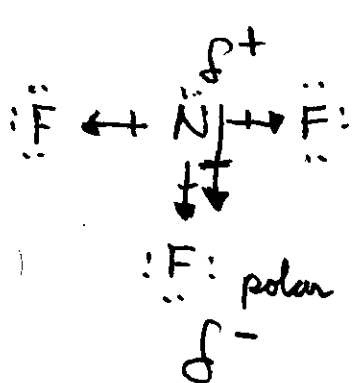
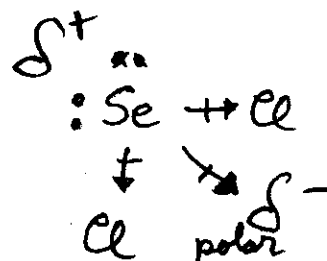
To the following stick structures add bond polarizations
 With consideration to symmetry (or asymmetry)
 determine if the molecule is polar or non-polar



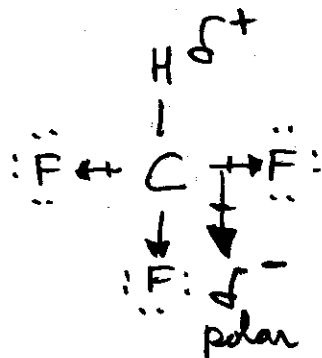
non-polar



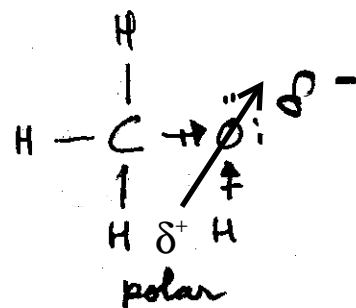
polar



polar
 δ^-



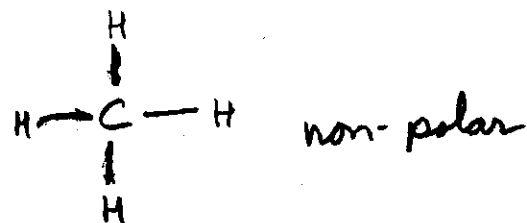
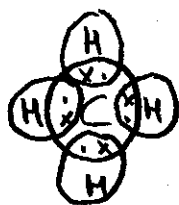
polar



polar

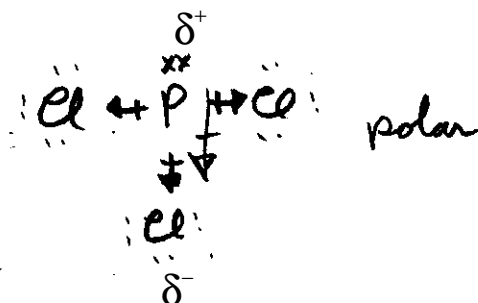
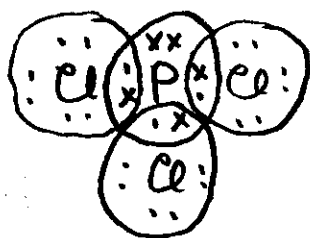
Write lewis structures, stick structures, bond polarizations
 and determine if the resulting molecule is polar or non-polar

C with H



non-polar

P with Cl



polar