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<u>SCH 3U Formula Quiz</u>

1. Given that a compound is 48.64 % carbon, 8.17 % hydrogen and 43.20 % oxygen by mass and that the mass of one mole of this compound is 296.32 g/mol, show a complete calculation for the molecular formula of this compound.

2. Perform a complete percentage my mass calculation for ammonium phosphate \rightarrow (NH₄)₃PO₄

3. Give an example of a chemical formula that is clearly a molecular formula and not an empirical formula. What is it about this formula that makes it impossible for it to represent a empirical formula.