Draw Lewis Dot Diagrams of the atoms, State how many electrons each atom needs, draw extra atoms where necessary to account for the composition of the covalent compounds, draw a Lewis Dot Diagram of the resulting covalent compound and finally, state how many electrons are shared in each bond and whether the bond is single, double or triple.

F F
CI
O O
N
H

(All of the above covalent compounds exist as gases at room temperature. Since each gas contains two atoms, they are referred to as diatomic gases. This is the naturally occurring state of each of these elements.)

N CI

CO

H O

H Cl

C H

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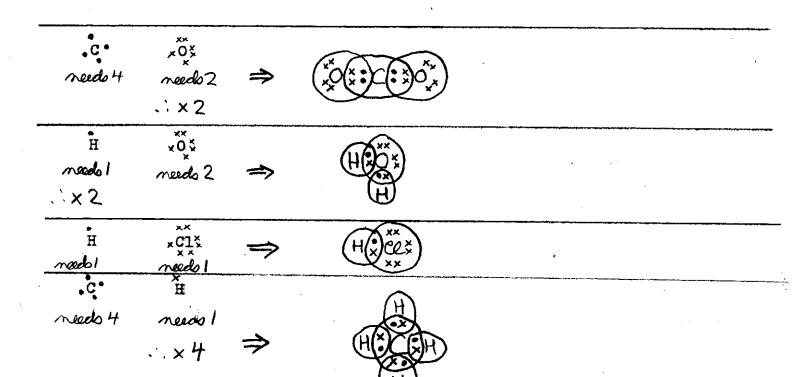
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F		⇒	FIRE	
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needo	2 needs 2	⇒		
.N needs	N.	⇒	(NEX) VE	
H needs	H	->	HRH	

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.F	needs!	⇒	FXFX	one single land	F2
.Cl2	*CT* nœds!	⇒	:celes	one single land	Clz
.0: needs2	xO\$ needs 2	⇒		one double bond	0,
.N° needs 3	N"	⇒	(X)	one triple bond	N ₂
H needs 1	# H needs l	⇒	HXH	one single bond	H ₂

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C. needo H	xox needo2 .; x 2	⇒	(**)(**)(**) two double bonds CO2
H macdol ×2	xŏ; meeds 2	⇒	HOSE two single londs H20
H nædol C	xC1X needs I	⇒	HOCEX one single bond Hel
needs 4	needs 1 ×4		HOCOH Jour single lands CH4