Exercise on Ionic Bonding

Draw Lewis Dot Diagrams of the neutral atoms, show the movement of electrons and draw Lewis Dot Diagrams of the resulting ions (include the charge on the ion). Where it is necessary, draw extra atoms to account for the law of constant composition of the ionic compounds

K Cl

Ca O

Sr F

Na O

K N

Mg Br

Ba Cl

Cs F

K Br

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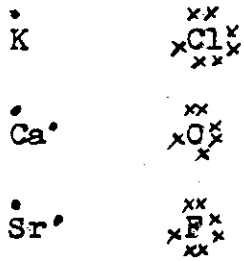

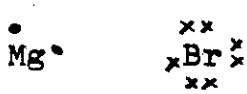

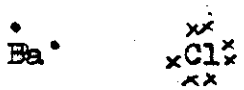

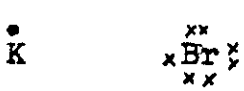

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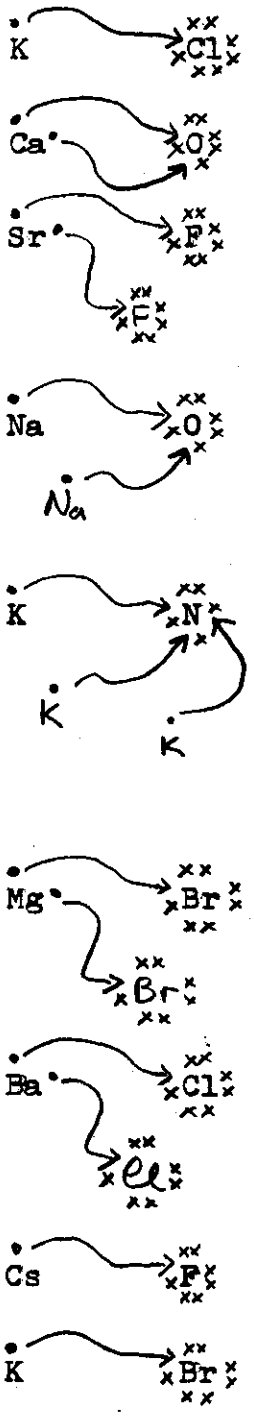

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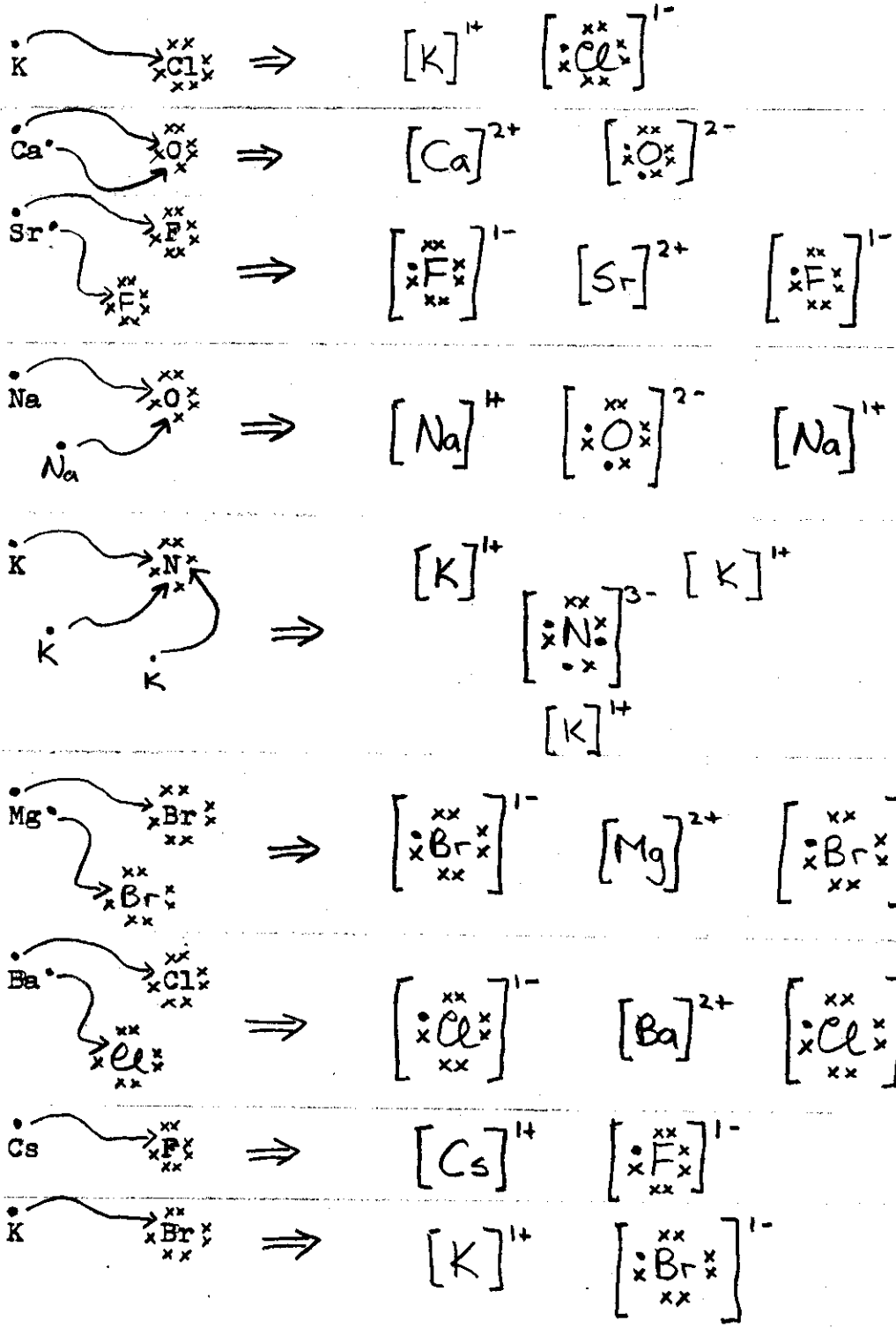

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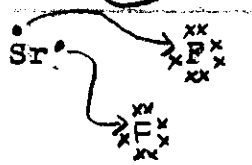
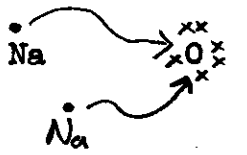
$\Rightarrow$ $[K]^{1+}$ $[:\ddot{Cl}:]^{1-}$ KCl

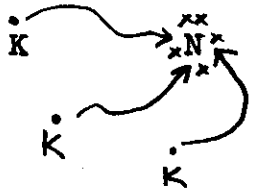
$\Rightarrow$ $[Ca]^{2+}$ $[:\ddot{O}:]^{2-}$ CaO



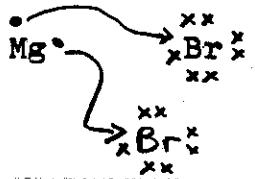
$\Rightarrow$ $[:\ddot{F}:]^{1-}$ $[Sr]^{2+}$ $[:\ddot{F}:]^{1-}$ SrF_2



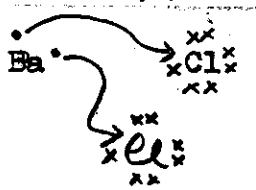
$\Rightarrow$ $[Na]^{1+}$ $[:\ddot{O}:]^{2-}$ $[Na]^{1+}$ Na_2O



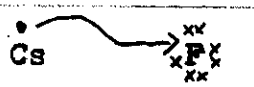
$\Rightarrow$ $[K]^{1+}$ $[:\ddot{N}:]^{3-}$ $[K]^{1+}$ $[K]^{1+}$ K_3N



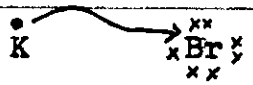
$\Rightarrow$ $[:\ddot{Br}:]^{1-}$ $[Mg]^{2+}$ $[:\ddot{Br}:]^{1-}$ $MgBr_2$



$\Rightarrow$ $[:\ddot{Cl}:]^{1-}$ $[Ba]^{2+}$ $[:\ddot{Cl}:]^{1-}$ $BaCl_2$



$\Rightarrow$ $[Cs]^{1+}$ $[:\ddot{F}:]^{1-}$ CsF



$\Rightarrow$ $[K]^{1+}$ $[:\ddot{Br}:]^{1-}$ KBr