

IONIC COMPOUNDS: Names and Formulas

1. Write the formulas for the following compounds.

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|------------------------------|----------------------------------|
| a. magnesium oxide _____ | k. tin (II) fluoride _____ |
| b. aluminum nitride _____ | l. lead (IV) nitride _____ |
| c. potassium sulfide _____ | m. iron (III) chloride _____ |
| d. calcium bromide _____ | n. copper (I) oxide _____ |
| e. aluminum sulfide _____ | o. antimony (III) sulfide _____ |
| f. beryllium oxide _____ | p. mercury (II) oxide _____ |
| g. strontium phosphide _____ | q. tin (IV) iodide _____ |
| h. sodium fluoride _____ | r. arsenic (III) phosphide _____ |
| i. lithium selenide _____ | s. cobalt (II) sulphide _____ |
| j. barium oxide _____ | t. tin (IV) sulphide _____ |

2. Write the names for the following compounds.

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|----------------------------------|--------------------------------|
| a. Li_2O _____ | k. PbS _____ |
| b. AlCl_3 _____ | l. SnO_2 _____ |
| c. MgS _____ | m. NiO _____ |
| d. CaF_2 _____ | n. CuI_2 _____ |
| e. Al_2O_3 _____ | o. PbCl_4 _____ |
| f. BeF_2 _____ | p. FeP _____ |
| g. K_3P _____ | q. AuBr_3 _____ |
| h. Mg_3P_2 _____ | r. Hg_2S _____ |
| i. CaO _____ | s. SbF_3 _____ |
| j. Ag_2S _____ | t. MnO_2 _____ |

* all require a roman numeral