

## Balancing Equations Race

- 1) \_\_\_ C<sub>3</sub>H<sub>8</sub> + \_\_\_ O<sub>2</sub> → \_\_\_ CO<sub>2</sub> + \_\_\_ H<sub>2</sub>O
- 2) \_\_\_ Al + \_\_\_ Fe<sub>3</sub>N<sub>2</sub> → \_\_\_ AlN + \_\_\_ Fe
- 3) \_\_\_ Na + \_\_\_ Cl<sub>2</sub> → \_\_\_ NaCl
- 4) \_\_\_ H<sub>2</sub>O<sub>2</sub> → \_\_\_ H<sub>2</sub>O + \_\_\_ O<sub>2</sub>
- 5) \_\_\_ C<sub>6</sub>H<sub>12</sub>O<sub>6</sub> + \_\_\_ O<sub>2</sub> → \_\_\_ H<sub>2</sub>O + \_\_\_ CO<sub>2</sub>
- 6) \_\_\_ H<sub>2</sub>O + \_\_\_ CO<sub>2</sub> → \_\_\_ C<sub>7</sub>H<sub>8</sub> + \_\_\_ O<sub>2</sub>
- 7) \_\_\_ NaClO<sub>3</sub> → \_\_\_ NaCl + \_\_\_ O<sub>2</sub>
- 8) \_\_\_ (NH<sub>4</sub>)<sub>3</sub>PO<sub>4</sub> + \_\_\_ Pb(NO<sub>3</sub>)<sub>4</sub> → \_\_\_ Pb<sub>3</sub>(PO<sub>4</sub>)<sub>4</sub> + \_\_\_ NH<sub>4</sub>NO<sub>3</sub>
- 9) \_\_\_ BF<sub>3</sub> + \_\_\_ Li<sub>2</sub>SO<sub>3</sub> → \_\_\_ B<sub>2</sub>(SO<sub>3</sub>)<sub>3</sub> + \_\_\_ LiF
- 10) \_\_\_ C<sub>7</sub>H<sub>17</sub> + \_\_\_ O<sub>2</sub> → \_\_\_ CO<sub>2</sub> + \_\_\_ H<sub>2</sub>O
- 11) \_\_\_ CaCO<sub>3</sub> + \_\_\_ H<sub>3</sub>PO<sub>4</sub> → \_\_\_ Ca<sub>3</sub>(PO<sub>4</sub>)<sub>2</sub> + \_\_\_ H<sub>2</sub>CO<sub>3</sub>
- 12) \_\_\_ Ag<sub>2</sub>S → \_\_\_ Ag + \_\_\_ S<sub>8</sub>
- 13) \_\_\_ KBr + \_\_\_ Fe(OH)<sub>3</sub> → \_\_\_ KOH + \_\_\_ FeBr<sub>3</sub>
- 14) \_\_\_ KNO<sub>3</sub> + \_\_\_ H<sub>2</sub>CO<sub>3</sub> → \_\_\_ K<sub>2</sub>CO<sub>3</sub> + \_\_\_ HNO<sub>3</sub>
- 15) \_\_\_ Pb(OH)<sub>4</sub> + \_\_\_ Cu<sub>2</sub>O → \_\_\_ PbO<sub>2</sub> + \_\_\_ CuOH
- 16) \_\_\_ Cr(NO<sub>2</sub>)<sub>2</sub> + \_\_\_ (NH<sub>4</sub>)<sub>2</sub>SO<sub>4</sub> → \_\_\_ CrSO<sub>4</sub> + \_\_\_ NH<sub>4</sub>NO<sub>2</sub>
- 17) \_\_\_ KOH + \_\_\_ Co<sub>3</sub>(PO<sub>4</sub>)<sub>2</sub> → \_\_\_ K<sub>3</sub>PO<sub>4</sub> + \_\_\_ Co(OH)<sub>2</sub>
- 18) \_\_\_ Sn(NO<sub>2</sub>)<sub>4</sub> + \_\_\_ Pt<sub>3</sub>N<sub>4</sub> → \_\_\_ Sn<sub>3</sub>N<sub>4</sub> + \_\_\_ Pt(NO<sub>2</sub>)<sub>4</sub>
- 19) \_\_\_ B<sub>2</sub>Br<sub>6</sub> + \_\_\_ HNO<sub>3</sub> → \_\_\_ B(NO<sub>3</sub>)<sub>3</sub> + \_\_\_ HBr
- 20) \_\_\_ ZnS + \_\_\_ AlP → \_\_\_ Zn<sub>3</sub>P<sub>2</sub> + \_\_\_ Al<sub>2</sub>S<sub>3</sub>